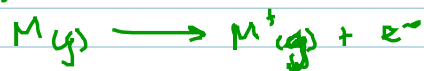
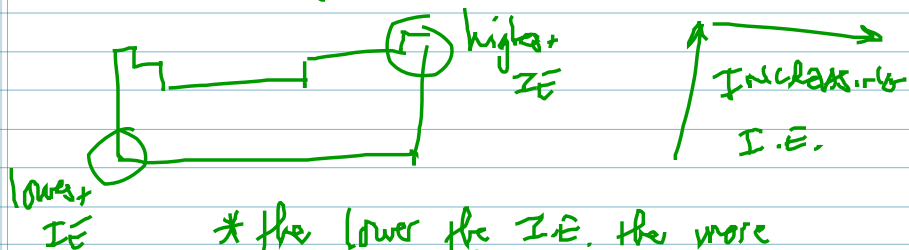


(3) IONIZATION ENERGY (IE)

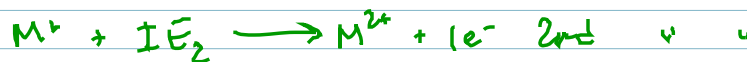
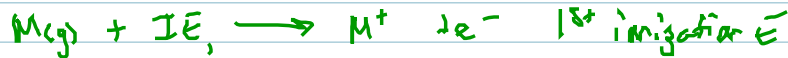
→ the minimum E needed to remove $1e^{-}$ from a gaseous state atom in the ground state



* a measure of how ~~#~~ tightly the e^{-} are being held



* the lower the IE, the more METALLIC the element



(4) ELECTRON AFFINITY (EA)

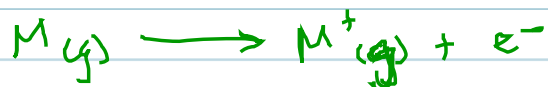
→ the negative of the ΔE when a gaseous atom accepts an e^{-}



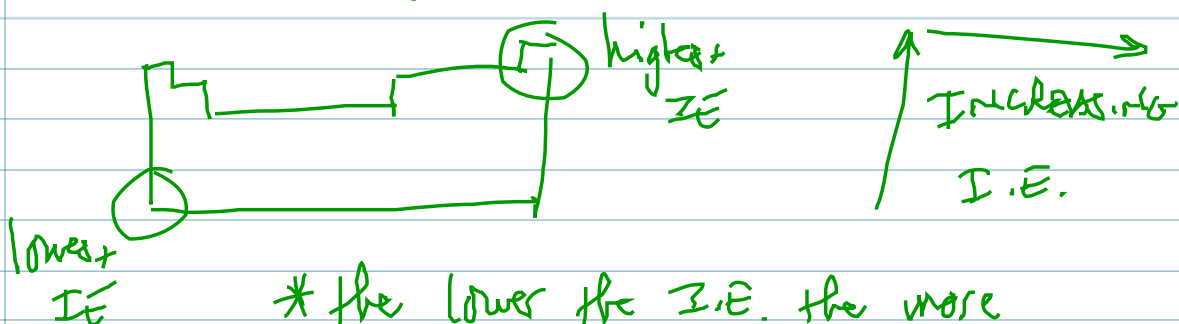
exothermic (-)

(3) IONIZATION ENERGY (IE)

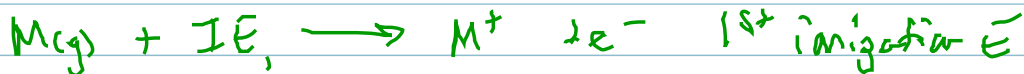
→ the minimum E needed to remove $1e^-$ from a gaseous state atom in the ground state



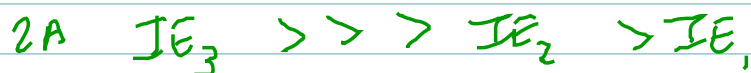
* a measure of how ~~st~~ tightly the e^- are being held



* the lower the I.E. the more METALLIC the element

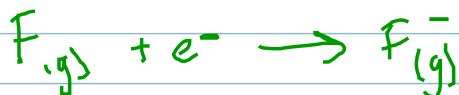


$$IE_1 < IE_2$$



(4) ELECTRON AFFINITY (EA)

→ the negative of the ΔE when a gaseous atom ACCEPTS an e^-



exothermic (-)

irregularities lower than expected EA values

2A

5A

) increasing e⁻/e⁻ repulsion